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Hydroplumboelsmoreite, $(\text{Pb}_{1-x}\text{Fe}_x)_2(\text{W}_{1.33}\text{Fe}_{3+0.67})_2\text{O}_6(\text{H}_2\text{O})$, A Redefined Mineral Species of the Elsmoreite Group from China

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Hydroplumboelsmoreite, $(\text{Pb}_1\Box_1)_{\Sigma 2}(\text{W}_{1.33}\text{Fe}_{0.67}^{3+})_{\Sigma 2}\text{O}_6(\text{H}_2\text{O})$, a redefined mineral species of the elsmoreite group from China

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Abstract

Hydroplumboelsmoreite (IMA21-C), $(\text{Pb},\Box)_2(\text{W},\text{Fe}^{3+})_2\text{O}_6(\text{H}_2\text{O})$, is a redefined elsmoreite-group mineral in the pyrochlore supergroup. It was found in a ‘jixianite’ cotype specimen provided by Mr. Liu Jianchang, who first found ‘jixianite’ in 1979 in the Jizhou District, Tianjin City, China. The mineral occurs as yellow to reddish brown aggregates, together with raspite and another elsmoreite-group mineral under study. Hydroplumboelsmoreite occurs in cryptocrystalline form and occasionally in octahedral microcrystalline form (under 20 μm in size). The crystals are colourless and translucent with a white streak, and the lustre is adamantine to greasy. Hydroplumboelsmoreite is isotropic, with a calculated refractive index of 2.29, a Mohs hardness of $\sim 4\frac{1}{2}$ –5, and a calculated density of 7.47 $\text{g}\cdot\text{cm}^{-3}$. The strongest five powder X-ray diffraction lines [d in $\text{\AA}(I)(hkl)$] are 6.070(28)(111), 3.012(100)(222), 2.603(32)(004), 1.836(35)(044) and 1.568(30)(226). The crystal structure was refined to $R_1 = 0.0459$ using 80 unique reflections collected with $\text{MoK}\alpha$ radiation, and the results show that the mineral is cubic, space group $Fd\bar{3}m$, with $a = 10.3377(5)$ \AA , $V = 1104.77(16)$ \AA^3 and $Z = 8$. Electron microprobe analyses and crystal structure refinement were used to determine the empirical formula: $(\text{Pb}_{1.05}\text{Sr}_{0.05}\text{Ce}_{0.07}^{3+}\text{Na}_{0.01}\Box_{0.82})_{\Sigma 2.00}(\text{W}_{1.32}\text{Fe}_{0.67}^{3+}\text{Zr}_{0.01})_{\Sigma 2.00}\text{O}_6[(\text{H}_2\text{O})_{0.43}\text{O}_{0.19}\Box_{0.38}]_{\Sigma 1.00}$. The mineral was named hydroplumboelsmoreite based on the predominance of Pb, W, and molecular H_2O in the A, B and Y sites, respectively.

Keywords: hydroplumboelsmoreite, elsmoreite, pyrochlore supergroup, crystal structure, tungsten

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Introduction

‘Jixianite’ used to be an approved mineral found in the Yanhe tungsten deposit in Jixian County, Hebei (currently Jizhou District in Tianjin City), China, by Liu (1979). It was an oxide containing W and Pb components and was approved by the International Mineralogical Association Commission on New Minerals and Mineral Names (IMA–CNMNC) at that time. However, because of the mineral’s tiny crystal size, the single crystal structure data could not be obtained. According to the powder-diffraction data, the diffraction characteristics are similar to those of pyrochlore-super group minerals, so it was classified as a pyrochlore-group mineral containing W and Pb, with the formula $\text{Pb}_2(\text{W},\text{Fe})_2(\text{O},\text{OH})_7$.

The general formula for the pyrochlore supergroup is $A_{2-m}B_2X_{6-w}Y_{1-n}$ (Atencio *et al.*, 2010). In 2010, based on the proposed new scheme for the nomenclature of the pyrochlore supergroup, which was approved by the IMA–CNMNC, ‘jixianite’ was discredited due to the absence of a detailed crystal structure study and lack of information about the Y site dominant anion (Atencio

et al., 2010). More recently, the status of the mineral’s name has been changed from ‘discredited’ to ‘questionable’ pending further research (Christy and Atencio, 2013).

Recently, we obtained a cotype specimen from Mr. Liu Jianchang who first found ‘jixianite’ in 1979 and we also collected new samples from the type locality. We selected 156 fragments from the cotype specimen to make an epoxy block for single-side polishing, including the residual part of the fragment used for structural data collection. Infrared spectroscopy, Raman spectroscopy and electron microprobe analysis were all performed on the polished surface of this sample group. Through this investigation, we discovered two members of the elsmoreite group within the pyrochlore supergroup in the cotype specimen of ‘jixianite’, including hydroplumboelsmoreite (IMA21-C, Miyawaki, 2021), which replaces ‘jixianite’, and a potential new mineral provisionally named ‘hydroelsmoreite’ (IMA2019-066, pending resubmission).

In this study, a 1.2 kW high-power microfocus rotating anode X-ray source and a hybrid pixel array detector single-crystal diffractometer were used to analyse the structure, and the single-crystal diffraction data were obtained for a small crystal (10 $\mu\text{m} \times 10 \mu\text{m} \times 10 \mu\text{m}$ in size). The results show that hydroplumboelsmoreite crystallises in the cubic system, space group $Fd\bar{3}m$, with lattice parameters of $a = 10.3377(5)$ \AA , $V = 1104.77(16)$ \AA^3 and $Z = 8$, so it has a pyrochlore type structure.

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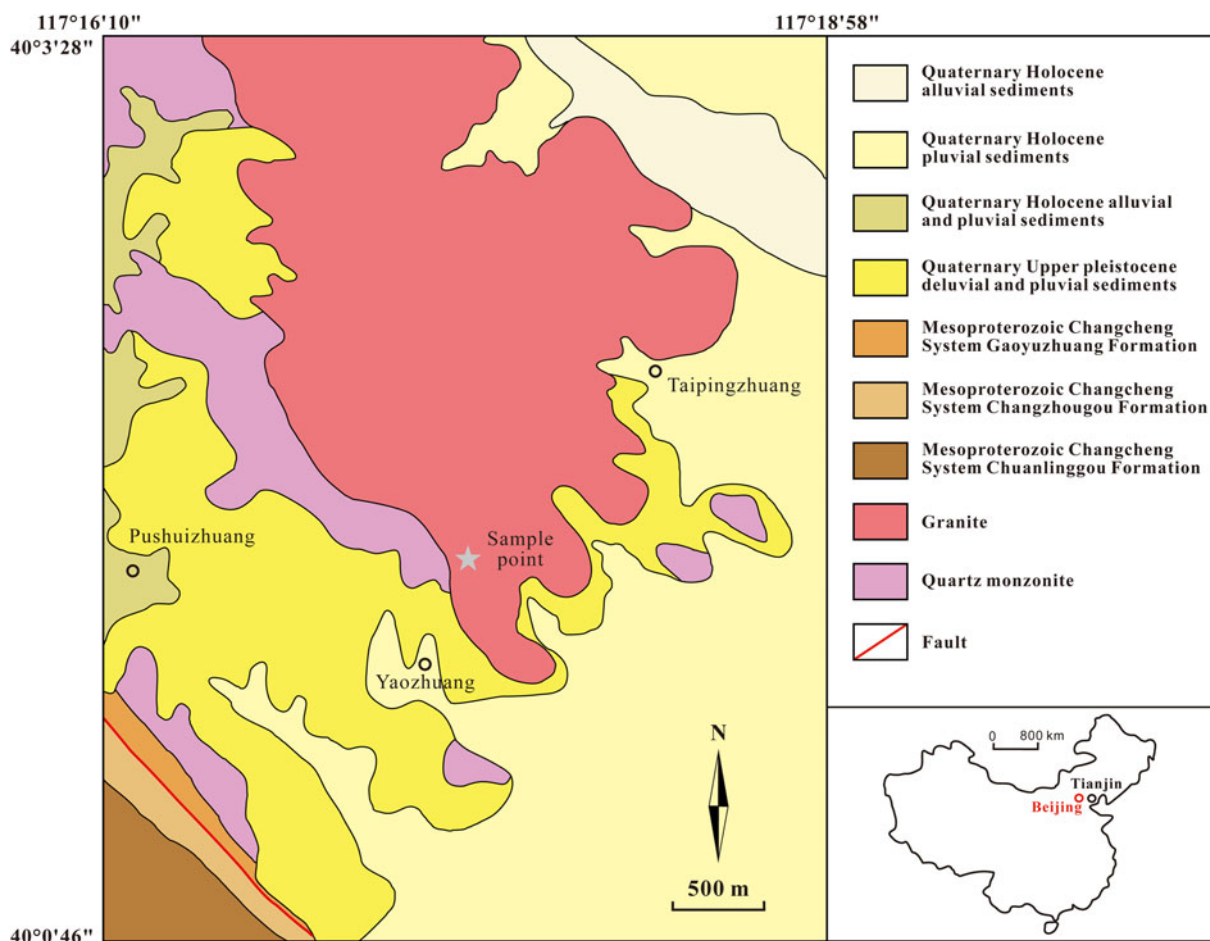


Fig. 1. Geological map of the locality.

The new electron-microprobe analyses carried out on hydroplumboelsmoreite are more accurate than the previous analyses, and were used to determine a simplified formula: $(\text{Pb}, \square)_2(\text{W}, \text{Fe}^{3+})_2\text{O}_6(\text{H}_2\text{O})$. According to the nomenclature of the pyrochlore supergroup, the mineral 'jixianite' is redefined as hydroplumboelsmoreite as it corresponds to a member of the elsmoreite group with Pb dominant in the A site and H_2O dominant in the Y site.

Hydroplumboelsmoreite and its name have been approved by the IMA–CNMNC and replaces 'jixianite'. The type specimen of hydroplumboelsmoreite has been registered within the collections of the Geological Museum of China (Xisi Yangrou hutong No. 15, Xicheng District, Beijing), China, under catalogue number M16125.

Occurrence and paragenesis

The mineral was discovered in the oxidised section of hydrothermal tungsten-bearing lead deposits in the Yanhe tungsten deposit in Jixian County, Hebei Province, China, which is now called the Jizhou District in Tianjin City ($40^\circ 1' 54.8'' \text{N}$, $117^\circ 17' 22.4'' \text{E}$).

The Jixian Yanhe tungsten deposit is situated in the middle of the Yanshan subsidence belt, the west wing of the Malanyu anticline, the southwest section of the Panshan short-axis anticline, and the southern border of the contact zone of the Panshan rock body (porphyritic quartz monzonite) (Fig. 1). The deposit is a high-temperature hydrothermal tungsten-bearing quartz vein type

deposit (Wang *et al.*, 2013). The tungsten mineralisation was related closely to magmatic activity and occurs in the inner contact zone of the rock. The metallogenic rock is mainly Late Indosinian intermediate-felsic rock, and the ore-bearing surrounding rocks are quartz monzonite and medium-to-fine-grained granite (Wen *et al.*, 2015). The occurrence, morphology, and scale of the tungsten-bearing quartz veins are strictly controlled by the lithology and structural fractures (Wen *et al.*, 2015). More than 290 quartz veins have been found in this area (Wang *et al.*, 2013).

Hydroplumboelsmoreite occurs in the oxidised section of the tungsten-bearing quartz veins (Liu, 1979). The primary minerals in the oxidised zone include quartz, pyrite, ferberite, cassiterite, chalcopyrite, scheelite, acanthite, silver and copper. The secondary minerals found nearby mainly include muscovite, bismutite, raspite, wulfenite, malachite, covellite, hydroplumboelsmoreite, and 'hydroelsmoreite' (under investigation).

Hydroplumboelsmoreite is an uncommon secondary mineral formed in the oxidised zone of the hydrothermal tungsten-bearing lead deposits. It mostly occurs as microcrystalline to cryptocrystalline aggregates, with an earthy, honeycombed, or crusty appearance (Fig. 2a), and the single crystals are tiny (Fig. 2b), exhibiting the characteristics of a hypergene mineral. The mineral is commonly intergrown with raspite (Fig. 2c,d) and 'hydroelsmoreite', and may be formed as a secondary mineral after scheelite, ferberite and wulfenite.

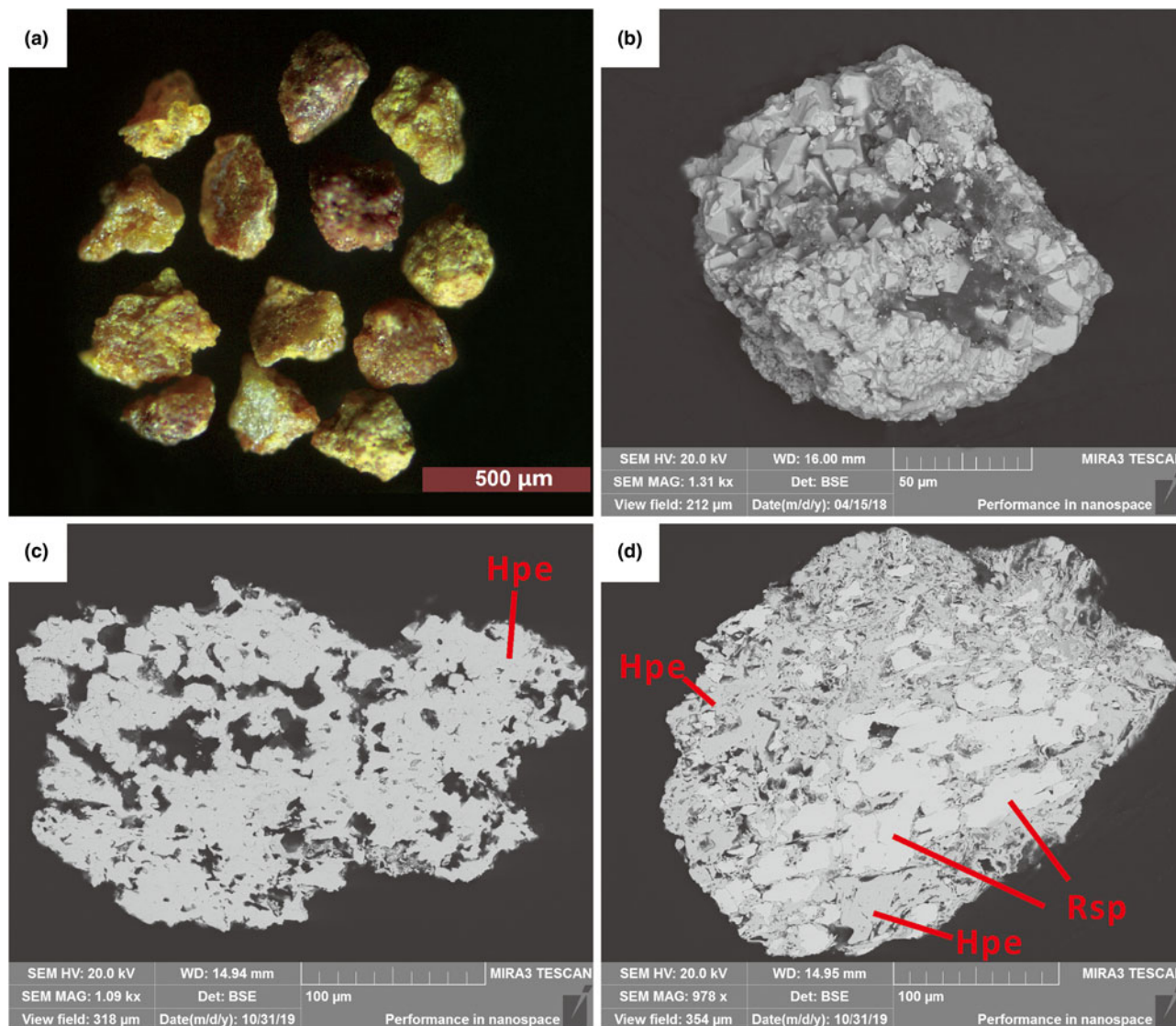


Fig. 2. Aggregate grains of hydroplumboelsmoreite: (a–c) aggregate grains of hydroplumboelsmoreite (Hpe) and (d) paragenetic raspite (Rsp) (symbols according to Warr, 2021).

Appearance and physical and optical properties

Generally, the aggregations reach 300 μm in size, but crystals are $<20 \mu\text{m}$ across (Fig. 2a,b). Monocrystals are rare, and they are sub-hedral or allotriomorphic granular, with the dominant form {111}. The aggregates are yellow or reddish brown (Fig. 2a). The crystals are colourless with a white streak, translucent with a greasy lustre, brittle with a conchoidal fracture, and no apparent twinning, cleavage, or parting was observed.

Hydroplumboelsmoreite is isotropic, and the refractive index calculated based on its chemical composition and Gladstone–Dale relationship ($N = K_d + 1$, K from Mandarino, 1981) is 2.29. No fluorescence was observed.

The density could not be measured because the single-crystal particles of this mineral are tiny and are generally intergrown with other minerals. However, the density calculated based on the empirical formula and the unit cell parameters refined from the single-crystal X-ray diffraction data is $7.47 \text{ g}\cdot\text{cm}^{-3}$. Vickers microhardness tests generally gave values of $278.59\text{--}302.3 \text{ kg}\cdot\text{mm}^{-2}$ (HV0.1kgf), and the Mohs hardness estimated by the VHN test is $\sim 4\frac{1}{2}\text{--}5$.

Spectroscopy

Infrared spectrum

The infrared (IR) spectrum of a polished surface of hydroplumboelsmoreite was obtained using a Bruker27 IR microscope equipped with a liquid nitrogen-cooled detector in reflection mode. The spectrum in the range of $4000\text{--}400 \text{ cm}^{-1}$ was collected by averaging 32 scans at a resolution of 4 cm^{-1} . The IR spectrum is shown in Fig. 3. The samples were stored at 110°C for 2 hours before the analysis to avoid the effect of adsorbed water on the test results. The 3552 cm^{-1} (strong) and 1631 cm^{-1} (medium) bands represent the stretching vibration and bending vibration of H_2O , respectively. The IR absorption related to the $\text{W}=\text{O}$ stretching vibration is represented by two peaks at 916 (weak) and 859 (weak) (Gadsden, 1975; Farmer, 1982). All of the other bands in the $400\text{--}700 \text{ cm}^{-1}$ range are due to the vibrations of the elsmoreite-type framework.

Raman spectrum

The Raman spectrum was obtained using a RENISHAW in Via-Reflex microscope Raman spectrometer equipped with a

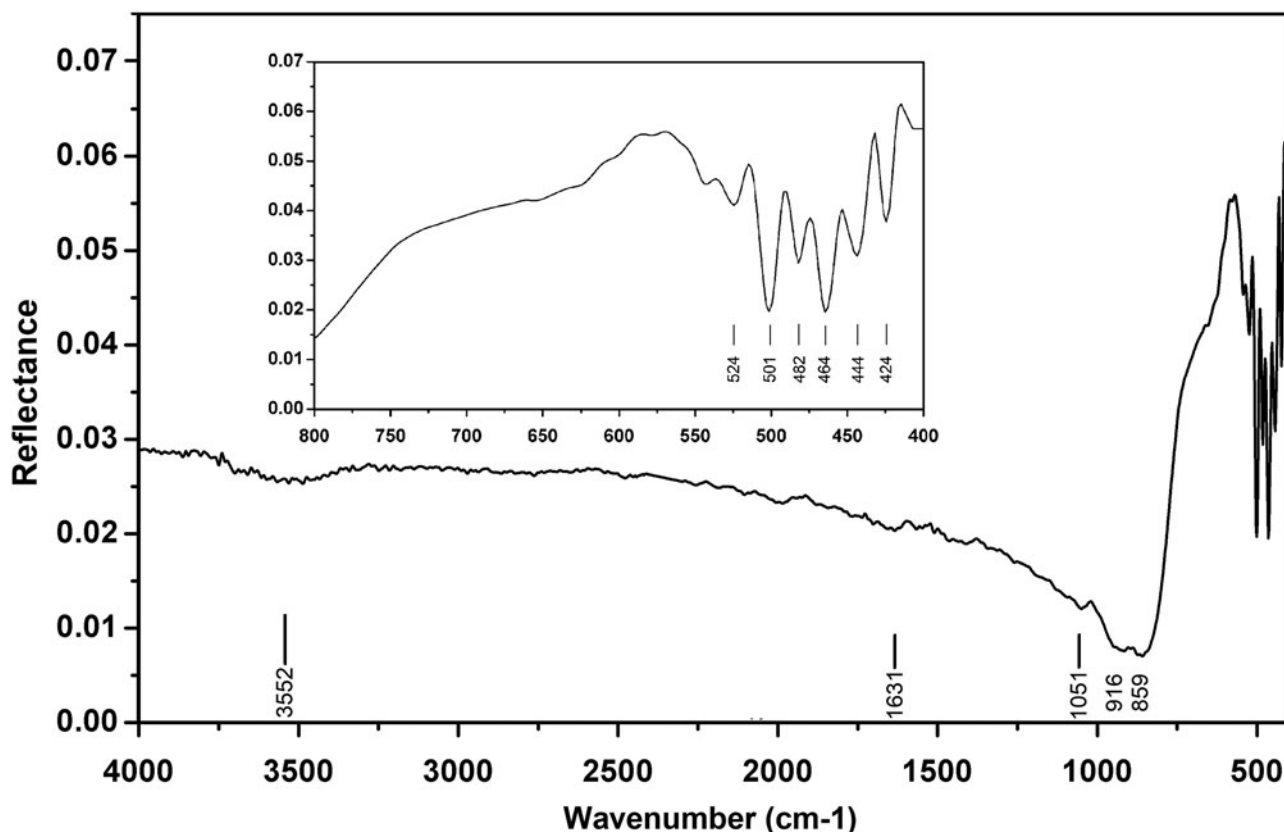


Fig. 3. Infrared spectrum of hydroplumboelsmoreite.

532 nm laser, in reflection mode, and the data acquisition time was 10 s. The Raman spectrum is shown in Fig. 4. The Raman spectrum of hydroplumboelsmoreite is characterised by an intense band at 891 cm^{-1} (W=O stretching) (Mills *et al.*, 2017). The bands at 695 and 501 cm^{-1} correspond to the O–W–O stretching modes, while the bands at 436 and 362 cm^{-1} correspond to the O–W–O bending modes (Gu *et al.*, 2006). The bands in the $100\text{--}350\text{ cm}^{-1}$ region are assigned to the lattice modes.

The band at 1583 cm^{-1} is an indicator of molecular H_2O , which is consistent with the IR spectrum. There is no obvious O–H stretching vibration band between 3400 and 3500 cm^{-1} .

Chemical composition

An average of ten spot-analyses was obtained using an electron microprobe (JEOL JXA-8100, wavelength-dispersive spectrometry mode (WDS), 20 kV, 10 nA and $2\text{ }\mu\text{m}$ beam diameter). The analytical data for hydroplumboelsmoreite and ‘hydroelsmoreite’, and the reported data for ‘jixianite’ are summarised in Table 1. It was found that the contents and total amounts of the different components of hydroplumboelsmoreite and ‘hydroelsmoreite’ are significantly different while the hydroplumboelsmoreite data are rather similar to those for ‘jixianite’.

The samples were stored again, at 110°C for 2 hours before the compositional analysis to avoid the effect of adsorbed water on the test results. Due to the very small particle-size of the hydroplumboelsmoreite, direct determination of H_2O was not possible. The H_2O was calculated based on the structure refinement. The presence of molecular H_2O and the absence of OH groups were confirmed by the infrared (IR) and Raman spectra. The dominant

occupant of the Y site (O2) was H_2O , which is consistent with the structure refinement, the bond valence sum, and the chemical composition analysis.

Given the known occurrence of hydroplumboelsmoreite in the oxidised zone, the appearance of typical supergene characteristics, the low scattering factor for the W1 site based on the structure refinement, and based on previous studies of elsmoreite-group minerals, all of the Fe was assumed to be trivalent and was assigned to the B site. Ce, Sr, Na and Ca were assumed to be confined to the A site.

On the basis of the results from the electron microscope analyses and crystal structure refinement (see below), the empirical formula is $(\text{Pb}_{1.05}\text{Sr}_{0.05}\text{Ce}_{0.07}\text{Na}_{0.01}\square_{0.82})_{\Sigma 2.00}(\text{W}_{1.32}\text{Fe}_{0.67}\text{Zr}_{0.01})_{\Sigma 2.00}\text{O}_6[(\text{H}_2\text{O})_{0.43}\text{O}_{0.19}\square_{0.38}]_{\Sigma 1.00}$, calculated on the basis of $B = 2$.

Moreover, because hydroplumboelsmoreite and ‘hydroelsmoreite’ are indistinguishable in terms of appearance, it is not possible to determine whether the chemical data for ‘jixianite’ (Liu, 1979) derived via wet-chemical analysis using larger amounts of subsample are accurate or not, and we speculate that the sample analysed may have been a mixture of raspite, hydroplumboelsmoreite and ‘hydroelsmoreite’.

X-ray diffraction and crystallography

Powder X-ray diffraction

The powder X-ray diffraction data (Table 2) were obtained using a Rigaku Oxford Diffraction XtaLAB PRO-007HF (MoK α , 50 kV and 24 mA) by the powder rotations method, from the $0.2\text{ mm} \times 0.2\text{ mm} \times 0.2\text{ mm}$ aggregated fragment used for all

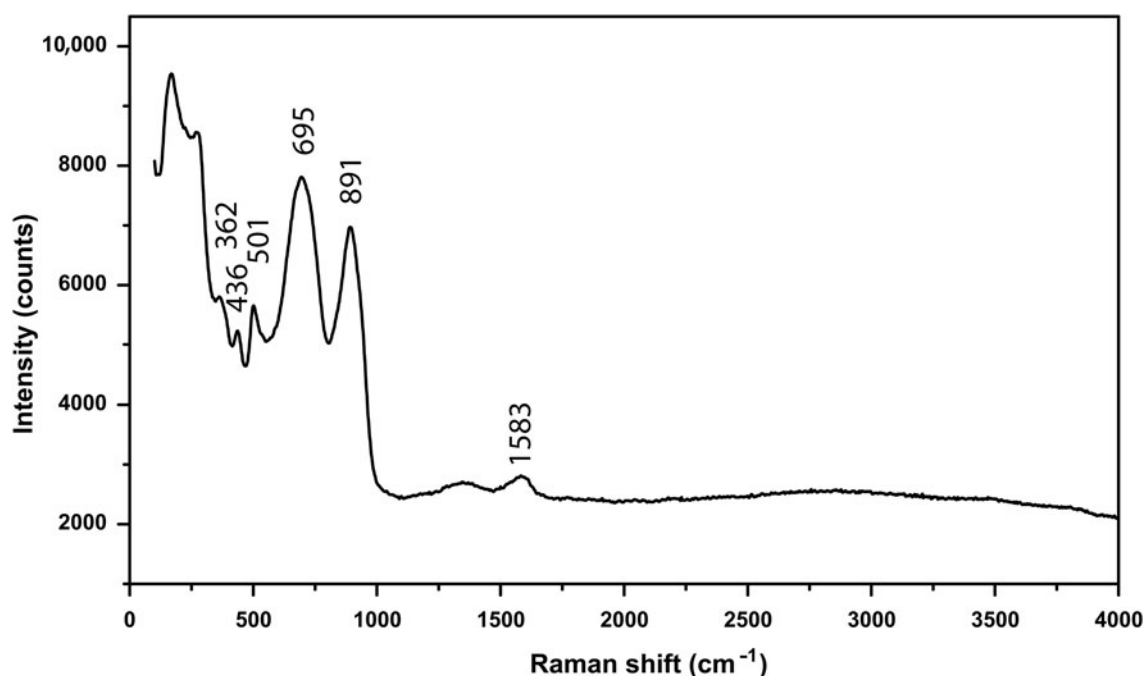


Fig. 4. Raman spectrum of hydroplumboelsmoreite.

Table 1. Chemical data (in wt.%) for hydroplumboelsmoreite, 'hydroelsmoreite' and 'jixianite'.

Constituent	Hydroplumboelsmoreite			'Hydroelsmoreite'			Probe standard	'Jixianite' **
	Mean	Range	S.D.	Mean	Range	S.D.		Mean
MgO	-	-	-	0.01	0-0.03	0.01	Pyrope	0.20
Na ₂ O	0.03	n.d.-0.12	0.04	0.27	0.05-0.38	0.09	Albite	-
WO ₃	48.71	46.23-51.33	1.69	59.82	52.72-63.23	3.22	Scheelite	50.39
PbO	37.49	35.27-40.40	1.66	22.70	19.77-27.49	2.23	PbCrO ₄	38.72
SrO	0.90	0.81-0.99	0.05	0.95	0.88-1.08	0.06	Celestite	-
Al ₂ O ₃	0.01	n.d.-0.04	0.01	0.20	0-0.75	0.30	Y ₃ Al ₅ O ₁₂	-
Fe ₂ O ₃	8.50	8.05-9.03	0.36	6.75	4.53-11.11	2.25	Synthetic Fe	7.24(FeO***)
K ₂ O	0.01	n.d.-0.04	0.01	0.04	0-0.11	0.03	Sanidine	-
MnO	0.01	n.d.-0.04	0.02	0.04	0-0.11	0.04	MnSiO ₃	-
CaO	0.04	n.d.-0.15	0.05	0.43	0.1-0.62	0.16	CaSiO ₃	-
Ta ₂ O ₅	0.03	n.d.-0.13	0.04	0.09	0-0.40	0.13	Manganotantalite	-
ZrO ₂	0.26	0.16-0.36	0.08	0.27	0.09-0.39	0.08	ZrO ₂	-
Ce ₂ O ₃	1.73	1.61-1.85	0.08	1.67	1.18-2.08	0.29	Monazite-(Ce)	-
Nb ₂ O ₅	0.06	n.d.-0.26	0.09	0.09	0-0.18	0.07	KNbO ₃	-
UO ₂	0.01	n.d.-0.06	0.02	-	-	-	Metallic uranium	-
F	n.d.	n.d.	n.d.	n.d.	n.d.	n.d.	Fluorapatite	-
TiO ₂	0.03	n.d.-0.08	0.04	-	-	-	Rutile	-
MoO	-	-	-	-	-	-	-	0.01
CuO	-	-	-	-	-	-	-	0.27
Total	97.81	97.16-98.56	0.43	93.34	92.11-94.50	0.79	-	96.83
H ₂ O*	1.27							

Notes: 'hydroelsmoreite' and 'jixianite' are not valid minerals approved by the IMA (Pasero, 2021). S.D. = standard deviation; n.d. = not detected; '-' = not analysed

* H₂O calculated by structure refinement; **Data from Liu (1979); ***FeO = Fe₂O₃ + FeO

the test. The powder X-ray diffraction pattern of hydroplumboelsmoreite (MoK α) is shown in Fig. 5. The five strongest powder X-ray diffraction lines [d in Å(I)(hkl)] are 6.070(28)(111), 3.012(100)(222), 2.603(32)(004), 1.836(35)(044) and 1.568(30)(226). The unit cell parameters refined from the powder-diffraction data using the *ChekcCell* software (Jean and Bernard, 2001) are as follows: cubic, $Fd\bar{3}m$ (#227), $a = 10.388(11)$ Å, $V = 1121(3)$ Å³ and $Z = 8$.

Single-crystal X-ray diffraction and structure determination

The single-crystal X-ray analysis was carried out using a Rigaku Oxford Diffraction XtaLAB PRO-007HF rotating anode micro-focus X-ray source (1.2 kW, MoK α , $\lambda = 0.71073$ Å) and a hybrid pixel array detector single-crystal diffractometer. The analytical conditions were 50 kV, 24 mA and an exposure time of 100 s per frame. A single crystal particle (10 $\mu\text{m} \times 10 \mu\text{m} \times 10 \mu\text{m}$ in

Table 2. Powder X-ray diffraction data (d in Å) for hydroplumboelsmoreite.

I_{obs}	I_{calc}	d_{obs}	d_{calc}	hkl
28	10	6.070	5.998	111
19	5	3.156	3.132	113
100	100	3.012	2.999	222
32	34	2.603	2.597	004
8	2	2.000	1.999	115
35	35	1.836	1.837	044
5	2	1.753	1.756	135
30	31	1.568	1.566	226
9	8	1.499	1.500	444
6	1	1.455	1.455	117
6	1	1.349	1.353	355
6	4	1.298	1.299	008
11	9	1.192	1.192	266
7	7	1.161	1.162	048

The strongest lines are given in bold.

size) from the powder X-ray diffraction sample was analysed. The structure solution and refinement were performed using *SHELX 2014* (Sheldrick, 2015), and the crystal structure was determined using the direct method. The details of the data collection and structure refinement for hydroplumboelsmoreite are summarised in Table 3. The atom coordinates and site occupancies are listed in Table 4. The anisotropic displacement parameters are listed in Table 5. Selected geometric parameters are listed in Table 6. The bond valences and bond-valence sums are presented in Table 7. The crystal structure of hydroplumboelsmoreite is shown in Fig. 6. The crystal structure refinement converged to $R_1 = 0.0459$ for 80 reflections with $|F_o| > 4\sigma F$ and 0.0509 for all 94 reflections. The crystallographic information file has been

deposited with the Principal Editor of *Mineralogical Magazine* and is available as Supplementary material (see below).

The research results show that this mineral crystallises in the cubic system, space group $Fd\bar{3}m$, with lattice parameters of $a = 10.3377(5)$ Å, $V = 1104.77(16)$ Å³ and $Z = 8$. The crystal structure is the pyrochlore type structure. W is [6]-coordinated (site 16c), is located at the centre of octahedron, and the W–O bond length is 1.944 Å. Pb, as a large cation, is [8]-coordinated (site 16d), is distributed in the channel, and the Pb–O bond length is 2.238 Å and 2.652 Å. Pb dominantly occupies the A site, while H₂O predominantly occupies the Y site (site 8b). Thus, hydroplumboelsmoreite is still a pyrochlore-supergrout mineral. The ideal end-member formula is $(\text{Pb}_1\Box_1)_{\Sigma 2}(\text{W}_{1.33}\text{Fe}_{0.67}^{3+})_{\Sigma 2}\text{O}_6(\text{H}_2\text{O})$ based on the crystal structure.

The structure refinement process and water content analysis were as follows: The occupancy of the A and B sites were fixed based on the compositional data collected from the homologous particle. The A site was occupied predominantly by Pb; and the B site was occupied by W and Fe. As other members of the elsmoreite group, i.e. hydroxyknoelsmoreite (Mills *et al.*, 2017), are found in lower symmetric space groups due to the ordered distributions of Fe and W, we tried to find lower symmetric space groups to solve the structure of hydroplumboelsmoreite, however no satisfactory results were found. Hence, the W and Fe should be disordered in our sample. The B site was fixed with $(\text{W}_{1.33}\text{Fe}_{0.67}^{3+})$ in the refinement process. The X site was assumed to be fully occupied, which was proven by the bond valence sums (BVS).

When we assumed the Y site (O2) was fully occupied by O[#] (e.g. O²⁻, OH⁻, or H₂O), the U_{eq} of O2 was 0.14766 Å², which seemed too high for O atoms. This may have been due to the following two reasons: (1) the quality of the diffraction data was not

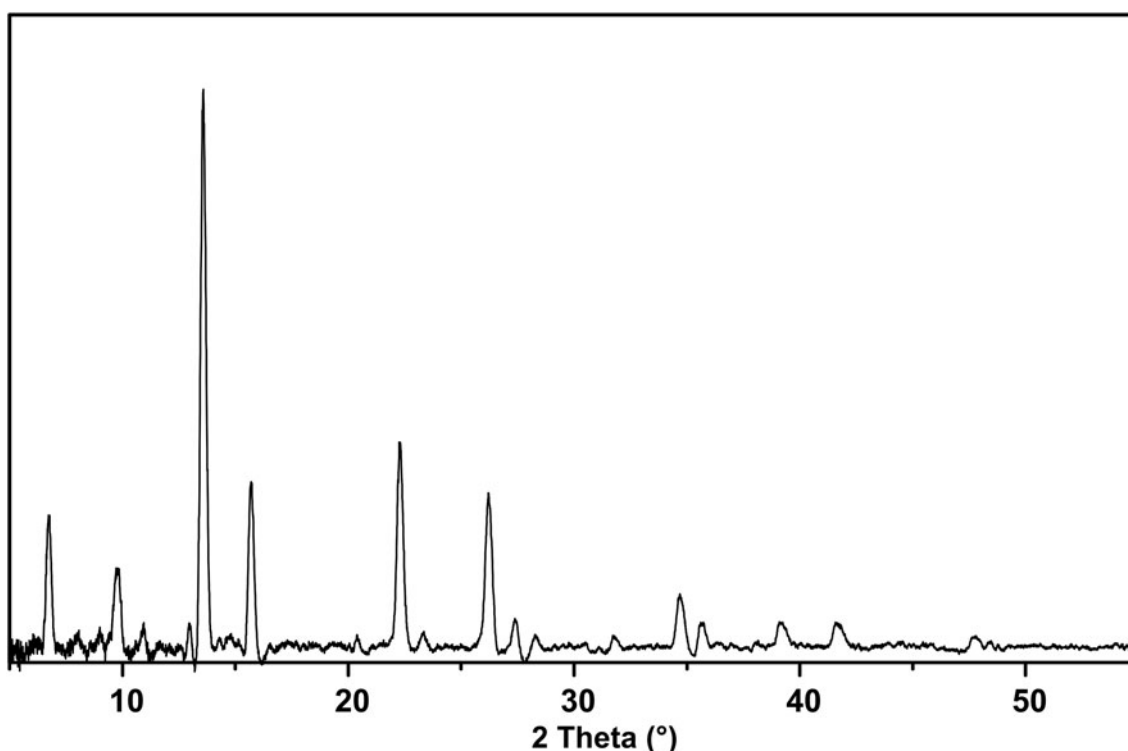


Fig. 5. Powder X-ray diffraction pattern of hydroplumboelsmoreite (MoK α).

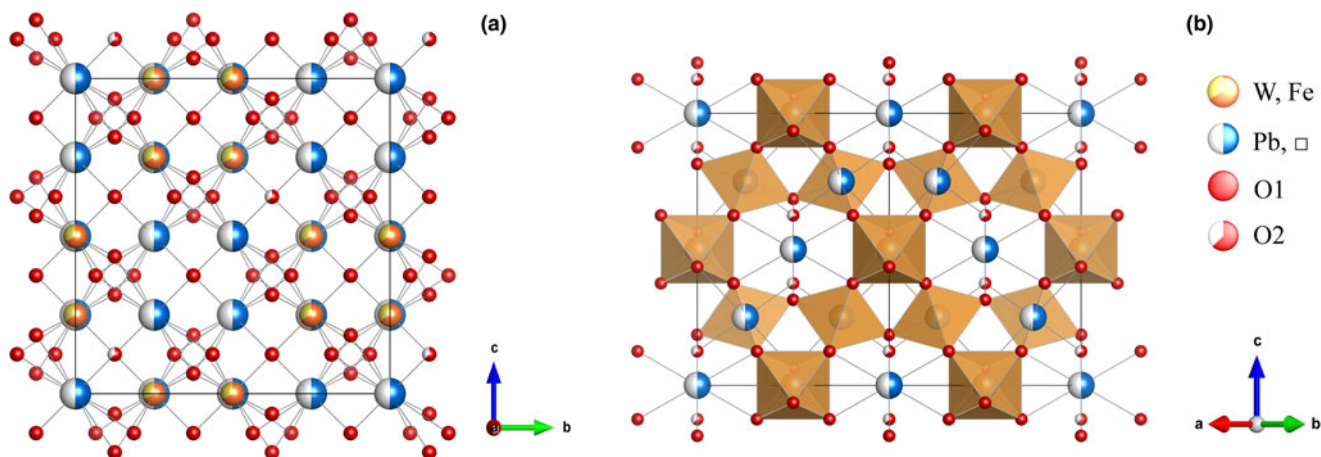


Fig. 6. General projection of the crystal structure of hydroplumboelsmoreite (a) along the a -axis and (b) the $[110]$ projection. A site: (Pb, \square). Drawn using Vesta (Momma and Izumi, 2011).

Table 8. Site occupancy factors (s.o.f.), R_1 and difference-Fourier values for the Y site under different U_{iso} values.

	U_{eq}	R_1	wR_2	s.o.f.	Difference Fourier
O2 atom omitted		0.0514	0.1600	0	+4.32 e^-
Full s.o.f.	0.14766	0.0464	0.1279	1	-0.24 e^-
Variable s.o.f and U_{iso}	0.08736	0.0456	0.1244	0.699	-0.52 e^-
Fixed U_{iso} 0.03	0.03	0.0467	0.1312	0.504	-1.01 e^-
Fixed U_{iso} 0.04	0.04	0.0465	0.1295	0.547	-0.88 e^-
Fixed U_{iso} 0.05	0.05	0.0461	0.1281	0.586	-0.77 e^-
Fixed U_{iso} 0.06	0.06	0.0459	0.1268	0.620	-0.69 e^-

such as $[111]^{\wedge}[1\bar{1}\bar{1}]$ and $[100]^{\wedge}[1\bar{1}\bar{1}]$ are 109.5° and 125.3° , respectively. These are broadly compatible with the 104° typical for an H_2O molecule and suggest possible orientations for that molecule. The Raman spectrum and BVS calculation also demonstrate

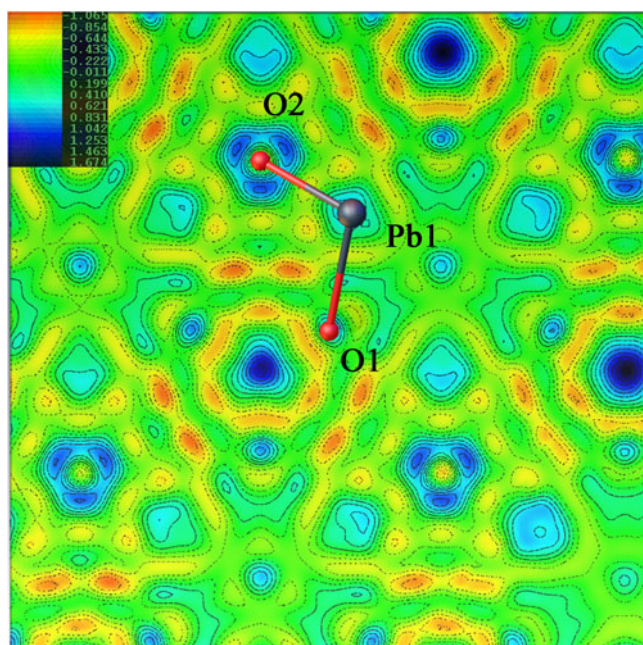


Fig. 7. Difference-Fourier electron density mapping along (111). Drawn using Olex2 (Dolomanov *et al.*, 2009).

the existence of molecular H_2O in the Y site. As the determination of the exact H positions using only the DFED is not reliable, we do not list the atomic coordinates of H. (5) Subsequently, ($Y_{O_{0.62}, \square_{0.38}}$) in the Y site is indicated, where $O^{\#}$ denotes the amount of occupation needed to balance the formula's charge recalculated into O atoms. It should be noted that the empirical formula from the EMPA shows the requirement of a charge of -0.38 in the Y site, thus $0.43H_2O$, $0.19O^{2-}$ and $0.38\square$ (vacancy) are assigned to the Y site. This corresponds to the 1.27 wt.% H_2O chemical composition and 99.08 wt.% total. Therefore, the empirical formula calculated on the basis of $B = 2$ is $(Pb_{1.05}Sr_{0.05}Ce_{0.07}Na_{0.01}\square_{0.82})_{\Sigma 2}(W_{1.32}Fe_{0.67}Zr_{0.01})_{\Sigma 2}O_6[(H_2O)_{0.43}O_{0.19}\square_{0.38}]_{\Sigma 1}$.

In fact, for the full or open occupancies of the Y site, or a fixed $U_{eq} = 0.06$, the refinements based on these three conditions lead to the Y site being occupied by $[(H_2O)_{0.81}O_{0.19}]$, $[(H_2O)_{0.51}O_{0.19}\square_{0.30}]$, or $[(H_2O)_{0.43}O_{0.19}\square_{0.38}]$, which all result in the conclusion that H_2O is dominant in the Y site.

Therefore, this new mineral is an elsmoreite-group mineral of the pyrochlore supergroup, and it was named hydroplumboelsmoreite according to the predominance of Pb in the A site and molecular H_2O in the Y site.

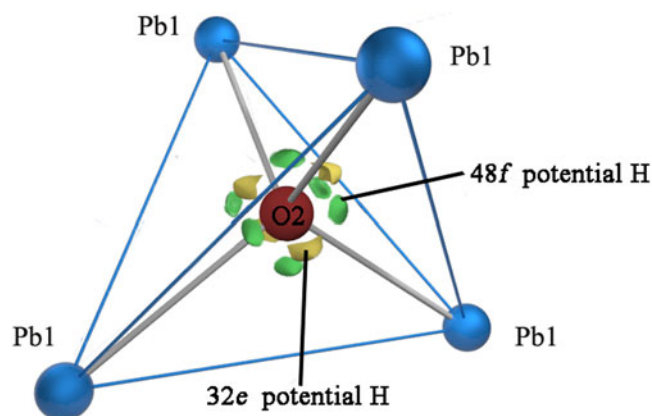


Fig. 8. Difference-Fourier electron density in the vicinity of the Pb1/O2 atoms. View of part of the hydroplumboelsmoreite structure, where the significant maxima of the positive difference electron density is located only around the O2 atom and four of them direct towards the face centres and six towards edge mid-points of the tetrahedron, away from the O2–Pb1 vectors.

Table 9. Comparison of the properties of the related elsmoreite-group minerals.*

Mineral species	Hydroplumboelsmoreite	'Hydroelsmoreite'	Hydrokenoelsmoreite-		Hydroxykenoelsmoreite
			3C	6R	
Simplified formula	$(\text{Pb}, \square)_2(\text{W}, \text{Fe}^{3+})_2\text{O}_6(\text{H}_2\text{O})$	$(\text{H}_2\text{O}, \text{Pb})_2(\text{W}, \text{Fe}^{3+})_2\text{O}_6(\text{H}_2\text{O})$	$(\square, \text{Na})_2(\text{W}, \text{Fe}^{3+}, \text{Al})_2(\text{O}, \text{OH})_6(\text{H}_2\text{O})$		$(\square, \text{Pb})_2(\text{W}, \text{Fe}^{3+}, \text{Al})_2(\text{O}, \text{OH})_6(\text{OH})$
Major component (wt.%)					
WO ₃	48.71	59.82	77.49	70.15	72.39
PbO	37.49	22.70	n.d.	n.d.	14.77
Al ₂ O ₃	0.01	0.02	2.41	0.97	1.67
Fe ₂ O ₃	8.50	6.75	5.83	8.17	5.99
Total (no H ₂ O)	97.81	93.34	88.68	82.68	95.01
H ₂ O	1.27 (calc.)	5.52 (calc.)	7.40 (calc.)	8.82 (calc.)	5.45 (calc.)
Total	99.08	98.86	96.08	90.85	100.46
Density (calc.)	7.47 g·cm ⁻³	6.78 g·cm ⁻³	6.025 g·cm ⁻³	6.025 g·cm ⁻³	5.806 g·cm ⁻³
Hardness					
– Mohs	4.5–5	3–4	±3	±3	~3
– Micro-indentation	278.59–302.3 kg·mm ⁻² (HV0.1kgf)	101.7–125.1 kg·mm ⁻² (HV0.1kgf)			
Unit cell parameters					
Crystal system, space group	cubic, <i>Fd</i> $\bar{3}m$	cubic, <i>Fd</i> $\bar{3}m$	cubic, <i>Fd</i> $\bar{3}m$	trigonal, <i>R</i> $\bar{3}$	trigonal, <i>R</i> $\bar{3}$
<i>a</i> (Å)	10.3377(5)	10.3421(5)	10.3065(3)	7.2882(2)	7.313(2)
<i>c</i> (Å)				35.7056(14)	17.863(7)
<i>V</i> (Å ³)	1104.77(16)	1106.18(17)	1094.80(6)	1642.51(9)	827(1)
<i>Z</i>	8	8	16	9	6
References	Our work	Our work	Mills <i>et al.</i> (2016); Williams <i>et al.</i> (2005)		Mills <i>et al.</i> (2017)
*Empirical formulae					
Hydroplumboelsmoreite	$(\text{Pb}_{1.05}\text{Sr}_{0.05}\text{Ce}_{0.07}\text{Na}_{0.01}\square_{0.82})_{\Sigma 2}(\text{W}_{1.32}\text{Fe}_{0.67}\text{Zr}_{0.01})_{\Sigma 2}\text{O}_6[(\text{H}_2\text{O})_{0.43}\text{O}_{0.19}\square_{0.38}]_{\Sigma 1}$				
'Hydroelsmoreite'	$[(\text{H}_2\text{O})_{0.76}\text{Pb}_{0.58}\text{Ce}_{0.06}\text{Sr}_{0.05}\text{Na}_{0.05}\text{Ca}_{0.04}\square_{0.46}]_{\Sigma 2}(\text{W}_{1.48}\text{Fe}_{0.49}\text{Al}_{0.02}\text{Zr}_{0.01})_{\Sigma 2}\text{O}_6(\text{H}_2\text{O}_{0.99}\text{O}_{0.01})_{\Sigma 1}$				
Hydrokenoelsmoreite-3C	$[\text{Na}_{0.28}\text{Ca}_{0.04}\text{K}_{0.02}(\text{H}_2\text{O})_{0.20}\square_{1.46}]_{\Sigma 2}(\text{W}_{1.47}\text{Fe}_{0.32}\text{Al}_{0.21}\text{As}_{0.01})_{\Sigma 2}(\text{O}_{4.79}(\text{OH})_{1.21})_{\Sigma 6}(\text{H}_2\text{O})$				
Hydrokenoelsmoreite-6R	$[\text{Na}_{0.24}\text{Ca}_{0.04}\text{K}_{0.03}(\text{H}_2\text{O})_{0.63}\square_{1.06}]_{\Sigma 2}(\text{W}_{1.42}\text{Fe}_{0.49}\text{Al}_{0.08}\text{As}_{0.01})_{\Sigma 2}(\text{O}_{4.65}(\text{OH})_{1.35})_{\Sigma 6}(\text{H}_2\text{O})$				
Hydroxykenoelsmoreite	$(\square_{1.668}\text{Pb}_{0.315}\text{Ca}_{0.009}\text{Na}_{0.005}\text{K}_{0.003}\text{Ba}_{0.001})_{\Sigma 2}(\text{W}_{1.487}\text{Fe}_{0.357}\text{Al}_{0.156})_{\Sigma 2}(\text{O}_{4.119}(\text{OH})_{1.881})_{\Sigma 6}(\text{OH})$				

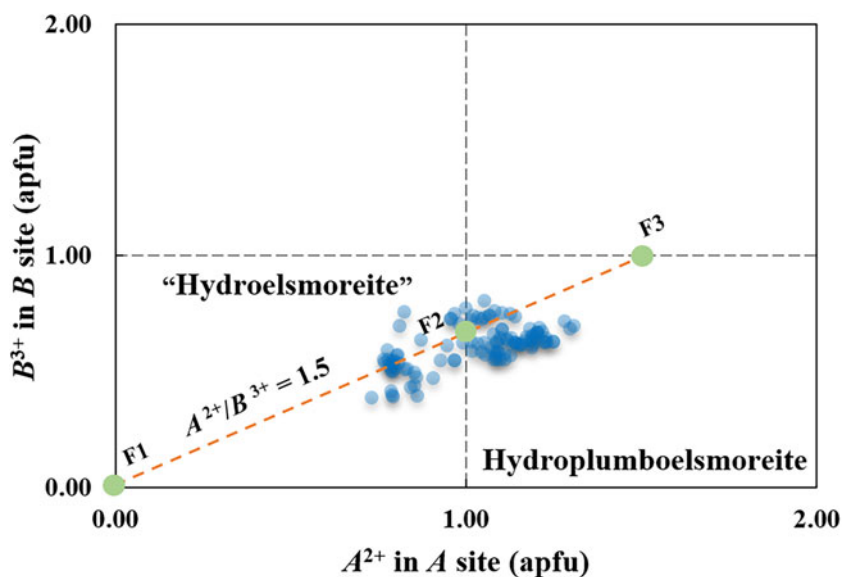


Fig. 9. Two-dimensional plots of the divalent A site cations and trivalent B site cations in hydroplumboelsmoreite and 'hydroelsmoreite'. F1 = $(A^0)_2W_2O_6(H_2O)$, F2 = $(PbA^0)_{\Sigma 2}(W_{1.33}Fe_{0.67}^{3+})_{\Sigma 2}O_6(H_2O)$, and F3 = $(Pb_{1.5}A^0_{0.5})_{\Sigma 2}(WFe^{3+})_{\Sigma 2}O_6(H_2O)$.

Discussion

Relationship to other mineral species

Hydroplumboelsmoreite is a new member of the pyrochlore supergroup. Its properties are compared with those of other members of the elsmoreite group in Table 9. A proposal for 'hydroelsmoreite' (IMA 2019-066) is being resubmitted to the CNMNC after revision. This phase is closely related to hydroplumboelsmoreite and was found in the same sample, but has obvious differences in chemical composition, crystal structure and hardness.

It should be noted that both of the previous two IMA-approved elsmoreite-group minerals, hydroxykenoelsmoreite (Mills *et al.*, 2016) and hydrokenoelsmoreite (Mills *et al.*, 2017), have been reported to deviate from the ideal cubic system due to ordering of Na and (\square, H_2O) in the A site or due to Fe^{3+} in the B site. However, the ordered distributions in these two sites were not observed in the hydroplumboelsmoreite analysed in this study and implies that the full cubic symmetry of the ideal pyrochlore structure is maintained. This phenomenon deserves more attention in elsmoreite-group minerals.

Chemical composition evolution

In the ideal formula, $(PbA^0)_{\Sigma 2}(W_{1.33}Fe_{0.67}^{3+})_{\Sigma 2}O_6(H_2O)$, which is charge balanced, the A site generally contains Pb^{2+} and other [8]-coordinated cations, and thus, the B site may contain subordinate Fe^{3+} or any other cations with lower charges to balance the charge. Neglecting minor substituents, the major components may have the following substitutions to make the formula electronically neutral: $(Pb_xA_{2-x}^0)_{\Sigma 2}(W_{2-y}Fe_y)_{\Sigma 2}O_6(H_2O)$. In addition, there is a proportionality between the bivalence cations in the A site and the trivalent cations in the B site, where $A^{2+}/B^{3+} = 1.5$ and $x = 1.5y$.

Thus, the electrically neutral end-member formula in the sense described by Hawthorne (2002) can be written as follows: if $y < 0.67$ and $x < 1$, the corresponding end-member formula is $(A^0)_2W_2O_6(H_2O)$ (labelled as F1 in Fig. 9), that is, 'hydroelsmoreite' or hydrokenoelsmoreite; otherwise, if $0.67 \leq y < 1$ and $1 \leq x \leq 2$, the corresponding end-member formula is $(PbA^0)_{\Sigma 2}(W_{1.33}Fe_{0.67}^{3+})_{\Sigma 2}O_6(H_2O)$ (labelled as F2 in Fig. 9) or

$(Pb_{1.5}A^0_{0.5})_{\Sigma 2}(WFe^{3+})_{\Sigma 2}O_6(H_2O)$ (labelled as F3 in Fig. 9), that is, hydroplumboelsmoreite. Hydroplumboelsmoreite and Pb-poor analogues such as 'hydroelsmoreite' and hydrokenoelsmoreite are named on the basis of dominant constituents of dominant valency groups in the Y , A and B sites in accord with Atencio *et al.* (2010). It should be noted that the end-member formula can be used as the ideal formula for hydroplumboelsmoreite is not unique although the former was found to be dominant in this study.

Making the assumption of negligible substitutions other than A^0 for Pb-dominant A^{2+} in the A site and Fe^{3+} for W^{6+} in the B site, 107 compositions obtained by WDS with polished fragments in the abovementioned epoxy block are plotted in Fig. 9. Although there is considerable solid solution that straddles the line between hydroplumboelsmoreite and 'hydroelsmoreite', the analysis points clearly cluster around the ideal composition F2 = $(PbA^0)_{\Sigma 2}(W_{1.33}Fe_{0.67}^{3+})_{\Sigma 2}O_6(H_2O)$.

Given the large charge difference between W and Fe, the preference for a W:Fe ratio of 2:1 may imply short-range order of these cations in the B site that is not reflected in the long-range average structure.

Conclusions

It appears that the original 'jixianite' sample with which the average chemical data (Liu, 1979) derived via wet-chemical analysis of powdered sample is a combination of raspite, hydroplumboelsmoreite and 'hydroelsmoreite'.

The compositional and crystal structural data collected in this study for a hydroplumboelsmoreite crystal unequivocally indicate that structurally, it belongs to the pyrochlore supergroup. Thus, it should be classified as hydroplumboelsmoreite in accordance with the classification rules for pyrochlore-supergroup minerals. It more specifically belongs to the elsmoreite group (W^{6+} predominantly in the B site), with Pb predominantly in the A site and H_2O predominantly in the Y site.

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Supplementary material. To view supplementary material for this article, please visit <https://doi.org/10.1180/mgm.2021.86>

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